SI session 13

March 6, 2014

1. Define a Buffer

2. Calculate the pH of a buffer solution of .1 M carbonic acid and .1 M of carbonate ion. *K*a of carbonic acid is 4.3 \* 10-7.

3. If the concentration of carbonic acid was .2 and .1 for carbonate ion, what is the pH of the solution.

4. When does pH=pKa?